

Prayas JEE 2026

Chemistry Redox Reaction

DPP: 1

- Q1** A compound contains **X**, **Y** and **Z** atoms. The oxidation states of **X** are **+a**, **Y** is **-b** and **Z** is **-c**. The molecular formula of the compound is $X_m Y_n Z_r$. Therefore, among the given relations, which one is correct?
 (A) $am + bn + cr = 0$
 (B) $am + bn = cr$
 (C) $am + cr = br$
 (D) $bn + cr = am$
- Q2** What is the oxidation state of **Xe** in Ba_2XeO_6 ?
 (A) 0 (B) +4
 (C) +6 (D) +8
- Q3** Which of the following have been arranged in the order of decreasing oxidation number of Sulphur?
 (A) $H_2S_2O_7 > Na_2S_4O_6 > Na_2S_2O_3 > S_8$
 (B) $SO^{2+} > SO_4^{2-} > SO_3^{2-} > HSO_4^-$
 (C) $H_2SO_5 > H_2SO_3 > SCl_2 > H_2S$
 (D) $H_2SO_4 > SO_2 > H_2S > H_2S_2O_8$
- Q4** Which of the following statements in true about oxidation state of **S** in $Na_2 S_4O_6$?
 (A) All **S**-atoms are in +2.5 state.
 (B) All **S**-atoms are in +2 state.
 (C) Two **S**-atoms are in 0 state and other two is in +5 state.
 (D) Two **S**-atoms are in -1 state and other two is in +6 state.
- Q5** The oxidation number of **Ba** in barium peroxide is:
 (A) +2 (B) -1
 (C) +4 (D) +6
- Q6** In nitric oxide (NO), the oxidation state of nitrogen is:
 (A) -2 (B) +1
 (C) -1 (D) +2
- Q7** The oxidation number of iodine in IF_5 is:
 (A) +5 (B) -5
 (C) -1 (D) +1
- Q8** Oxidation number of sulphur in Caro's acid is
 (A) +6 (B) +4
 (C) +8 (D) +7
- Q9** The oxidation number of sulphur in $H_2S_2O_8$ is
 (A) +2 (B) +6
 (C) +7 (D) +14
- Q10** The oxidation number of **Cl** in $KClO_3$ is:
 (A) +5 (B) -5
 (C) +3 (D) -3


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Answer Key

Q1 D
Q2 D
Q3 A, C
Q4 C
Q5 A

Q6 D
Q7 A
Q8 A
Q9 B
Q10 A



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