Prayas JEE 2026

Chemistry

Redox Reaction

DPP: 2

Q1 An oxide of metal have 20% oxygen, the equivalent weight of oxide is

(A) 32 (B) 48

(C) 40 (D) 52

Q2 The equivalent weight of $NaHC_2O_4$ in reaction with $NaOH$ is

(A) 112 (B) 56

(C) 224 (D) 84

Q3 H_2O_2 disproportionates into H_2O and O_2. The equivalent weight of H_2O_2 in this reaction is

(A) 34 (B) 17

(C) 68 (D) 8.5

Q4 The number of moles of $Cr_2O_7^{2-}$ needed to oxidize 0.136 equivalent of $N_2H_5^+$ through the reaction

$N_2H_5^+ + Cr_2O_7^{2-} \rightarrow N_2 + Cr^{3+} + H_2O$ is

(A) 0.023 (B) 0.091

(C) 0.136 (D) 0.816

Q5 $KMnO_4$ oxidises oxalic acid in acidic medium. The number of CO_2 molecule produced as per the balanced equation is

(A) 10 (B) 8

(C) 6 (D) 3

Q6 When N_2 is converted into NH_3, the n factor of nitrogen will be

(A) 1 (B) 2

(C) 3 (D) 6

Q7 If molecular weight of $KMnO_4$ is 'M', then its equivalent weight in acidic medium would be

(A) M

(B) $M/2$

(C) $M/5$

(D) $M/4$

Q8 In the conversion $NH_2OH \rightarrow N_2O$, the equivalent weight of NH_2OH will be

(A) $M/4$

(B) $M/2$

(C) $M/5$

(D) $M/1$

Answer Key

Q1	C	Q5	A
Q2	A	Q6	D
Q3	A	Q7	C
Q4	A	Q8	B

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